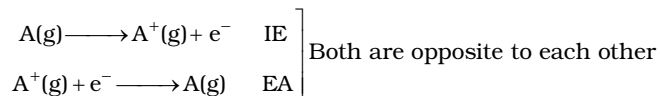


76.(A) I.E. of A = E.A. of A^+



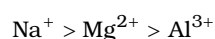
77.(AD) Ionization enthalpy is the energy required in joule to remove an electron and electron gain enthalpy is the energy released in joule to accept an electron.

Electronegativity is a relative term to attract the shared pair of electron and metallic character is the chemical property associated with metal that how easily it can lose their electron.

78.(C) Representative elements \Rightarrow elements of s & p block.

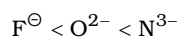
Write electronic configuration & if last e^- goes in s/p orbital. It is a representative element.

79.(C) Decreasing size of ion is gaseous state



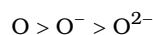
Higher is the positive charge, lower is the size in isoelectronic species

\Rightarrow Increasing size of ions in gaseous state



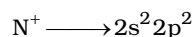
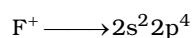
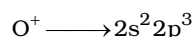
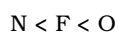
Higher is the negative charge, larger is the size in isoelectronic species

\Rightarrow Decreasing order of electron affinity



More is the charge, lower is the z_{eff} , lower is electron affinity.

\Rightarrow Increasing order of IE_2



Half filled electronic configuration is extra stable.

80.(A) $IE_{.3} > IE_{.2} > IE_{.1}$

81.(B) Along the period, I.E increases (generally) as Z_{eff} increases

Down the group, I.E. decreases as size increases

82. (C) I.E of $Li^+ > O^+$ due to more Z_{eff} in case of Li^+ .

83.(C) $\begin{array}{ccccc} Li & & Be & & B \\ & \searrow & & \searrow & \\ & Mg & & Al & \searrow \\ & & & & Si \end{array}$ Show diagonal similarity.

84.(ABC) $\Delta H_1, \Delta H_2$ and ΔH_3 are negative whereas ΔH_4 , is positive.

85.(A) Only atomic number 33 has unpaired 4p electrons.